

General Science Reading Log
Changes of State
p. 638-641

Name _____
 Date _____ Period _____

Part 1:

Before you read this section, read each statement below. If you believe a statement is true, place a check in the *Agree* column. If you believe a statement is false, place a check in the *Disagree* column. Be ready to explain your choices.

Agree

Disagree

- | | | |
|--------------------------|--------------------------|---|
| <input type="checkbox"/> | <input type="checkbox"/> | 1. Temperature is a measure of the speed (kinetic energy) of molecules. |
| <input type="checkbox"/> | <input type="checkbox"/> | 2. The melting point of water is the same as the freezing point. |
| <input type="checkbox"/> | <input type="checkbox"/> | 3. When a material is changing state, the temperature changes. |
| <input type="checkbox"/> | <input type="checkbox"/> | 4. Some materials don't melt: they change directly from a solid to a gas. |

Part 2:

Do the reading log on the back side of this page.

Part 3:

After you have done the reading log, look again at the questions and your answers in Part 1. If the information you read, supports your answer, check in the *Yes* column below, then write what the text says in column (A). If the information does not support your choices, place a check in the *No* column and write what the text says in column B.

***Text supports
my choice.***

	Yes	No	(A) Why is my choice correct?	(B) Why is my choice incorrect?
1.	<input type="checkbox"/>	<input type="checkbox"/>	_____	_____
			_____	_____
			_____	_____
2.	<input type="checkbox"/>	<input type="checkbox"/>	_____	_____
			_____	_____
			_____	_____
3.	<input type="checkbox"/>	<input type="checkbox"/>	_____	_____
			_____	_____
			_____	_____
4.	<input type="checkbox"/>	<input type="checkbox"/>	_____	_____
			_____	_____
			_____	_____

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Paragraph #	What is this about?	What is special?
p. 638 ¶ #1		
p. 638 ¶ #2 & 3		
p. 638 ¶ #4 <i>and</i> p. 639 ¶ #1		
p. 639 ¶ #2	kinetic energy & potential energy	
p. 639 ¶ #3 <i>and</i> p. 640 ¶ #1		
p. 640 ¶ #2	vaporization	
p. 640 ¶ #3	heating curve of water	Compare graph to yours. How is it similar?
p. 641 ¶ #1		

pages 638 - 641

Paragraph #	What is this about?	Why is it special?
p. 638 ¶ #1	particles	<ul style="list-style-type: none"> all matter made of particles dif. materials made of dif. types of particles particles are always moving closer particles have greater attractive forces
p. 638 ¶ #2 & 3	temperature	<ul style="list-style-type: none"> liquids expand when hot temperature is measure of speed of molecules temperature = average KE of molecules $KE = \frac{1}{2} mv^2$
p. 638 ¶ #4 and p. 639 ¶ #1	melting & freezing points	<ul style="list-style-type: none"> as temp increases, KE of particles increases normal melting pt. = temp when water changes from solid to liquid normal freezing point = temp when water changes from liquid to solid
p. 639 ¶ #2	kinetic energy & potential energy	<ul style="list-style-type: none"> when temp of water stayed same (0c), avg. KE did not change all energy caused phase change (change in PE) if change in KE, then change in temp.
p. 639 ¶ #3 and p. 640 ¶ #1	boiling point	<ul style="list-style-type: none"> 100°C = normal boiling point if atm. pressure < 1, water will boil < 100°C, like on Mt. Rainier
p. 640 ¶ #2	vaporization	<ul style="list-style-type: none"> at boiling pt, temp. didn't change even with heat energy added temp would stay same until liquid vaporized
p. 640 ¶ #3	heating curve of water	Compare graph to yours. How is it similar?
p. 641 ¶ #1	sublimation	<ul style="list-style-type: none"> dry ice (solid CO₂) doesn't have melting point dry ice has sublimation point sublimation = solid to gas